Correction to "Cold-trapped organic compounds at the poles of the Moon and Mercury: Implications for origins"

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[1] In the paper "Cold-trapped organic compounds at the poles of the Moon and Mercury: Implications for origins" by Jo Ann Zhang and David A. Paige (*Geophysical Research Letters*, *36*, L16203, doi:10.1029/2009GL038614), an incomplete version of Figure 1 was published. The complete Figure 1 and its caption appear here.

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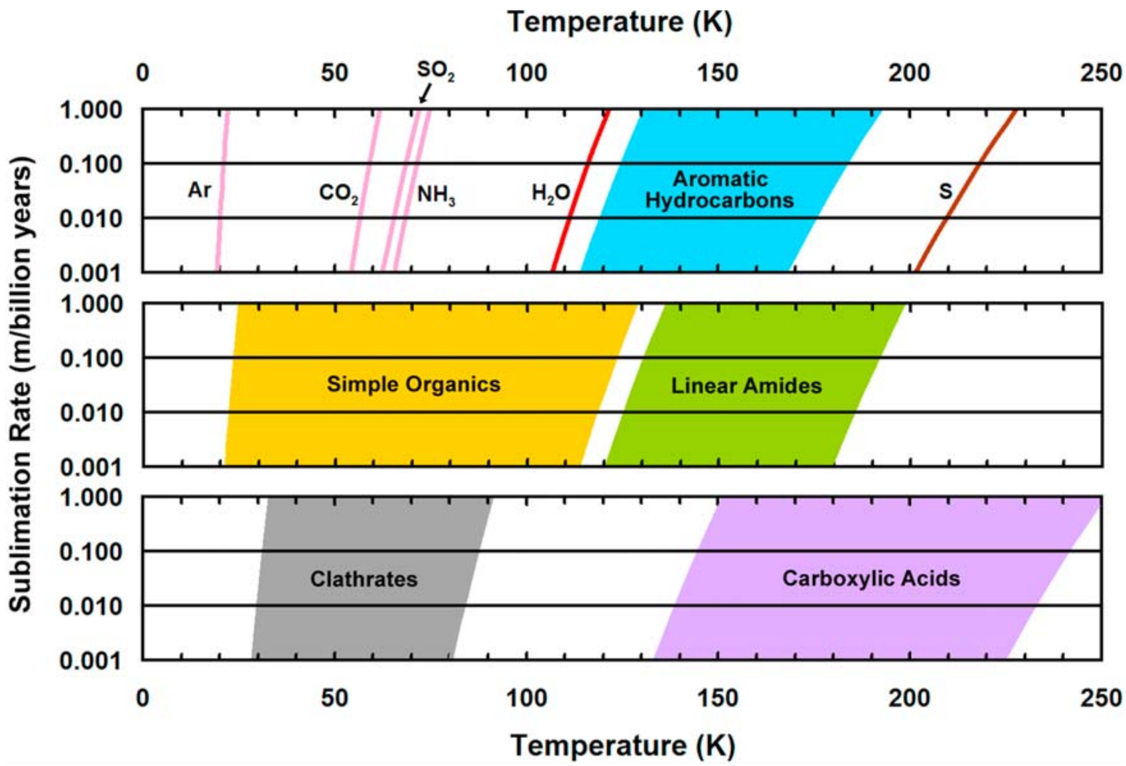

	Chemical Formula	Name	M (g/mol)	T_v (K)		Chemical Formula	Name	M (g/mol)	T_v (K)
Inorganics	N_2	Nitrogen	28.02	16.2	Aromatic Hydrocarbons	C_7H_6O	Benzaldehyde	106.13	114.6
	CO	Carbon monoxide	28.01	18.2		$C_7H_6O_2$	Salicylaldehyde	122.12	115.6
	Ar	Argon	39.95	19.5		C_6H_7N	Aniline	93.13	116.5
	Kr	Krypton	83.80	24.5		$C_6H_5NO_2$	Nitrobenzene	123.11	122.8
	Xe	Xeon	131.30	36.1		$C_{12}H_{10}$	Biphenyl	154.20	131.3
	H_2S	Hydrogen sulfide	34.09	50.6		$C_{10}H_8O$	1-Naphthol	144.16	137.7
	CO_2	Carbon dioxide	44.01	54.3		C_6H_6O	Phenol	94.11	141.9
	NH_3	Ammonia	17.03	65.5		$C_{10}H_8O$	2-Naphthol	144.16	139.1
	NH_4SH	Ammonium hydrosulfide	51.12	96.1		$C_{10}H_8$	Naphthalene	128.16	146.1
	SO_2	Sulfur dioxide	64.07	62.3		C_7H_8O	Benzyl alcohol	108.13	153.3
	H_2O	**Water**	18.02	106.6		$C_{14}H_{10}$	Phenanthrene	178.22	157.8
	S	**Sulfur**	32.07	201.5		$C_{12}H_{18}$	Hexamethylbenzene	162.26	163.2
Simple Organics	CH_4	Methane	16.04	22.0		$C_{14}H_{10}$	Anthracene	178.22	167.5
	OCS	Carbonyl sulfide	60.08	46.8	Linear Amides	C_3H_7NO	Dimethylformamide	70.10	120.7
	C_5H_{12}	Pentane	72.15	73.6		C_3H_7NO	Methylacetamide	73.10	120.8
	CS_2	Carbon disulfide	76.15	74.4		C_2H_5NO	Formamide	45.04	129.8
	HCN	Hydrogen cyanide	27.03	80.5		CH_3NO	Methylformamide	59.07	130.1
	C_7H_8	Toluene	92.13	87.6		C_2H_5NO	Acetamide	59.07	153.3
	$C_5H_{10}O$	3-Pentanone	96.21	92.8		C_4H_9NO	Dimethylacetamide	87.12	154.9
	NH_4CN	Ammonium cyanide	44.06	93.8		C_4H_9NO	Methylpropanamide	87.12	179.1
	$C_5H_{10}O_2$	Ethyl propanoate	102.13	103.6	Carboxylic Acids	$C_5H_{10}O_2$	Valeric acid	102.13	133.6
	$NH_4CO_2NH_2$	Ammonium carbonate	78.08	107.4		$C_6H_{12}O_2$	Caproic acid	116.16	148.4
	NH_4HCO_3	Ammonium bicarbonate	79.06	113.3		$C_7H_{14}O_2$	Enanthic acid	130.18	159.1
Clathrates	$Ar\cdot 6H_2O$	Argon clathrate	148.05	28.9		$C_8H_{16}O_2$	Caprylic acid	144.23	172.4
	$N_2\cdot 6H_2O$	Nitrogen clathrate	136.12	30.5		$C_9H_{18}O_2$	Pelargonic acid	158.23	181.1
	$O_2\cdot 6H_2O$	Oxygen clathrate	140.10	31.9		$C_{10}H_{20}O_2$	Capric acid	172.26	185.4
	$CO_2\cdot 7H_2O$	Carbon dioxide clathrate	170.12	33.4		$C_{12}H_{24}O_2$	Lauric acid	200.31	201.3
	$N_2\cdot 7H_2O$	Nitrogen clathrate	154.13	33.4		$C_{16}H_{32}O_2$	Palmitic acid	228.36	221.4
	$CH_4\cdot 7H_2O$	Methane clathrate	142.15	41.8		$C_{18}H_{36}O_2$	Stearic acid	256.42	224.3
	$CO_2\cdot 6H_2O$	Carbon dioxide clathrate	152.12	50.9	Fullerene	C_{60}	Fullerene	720.60	465.0
	$Xe\cdot 6H_2O$	Xeon clathrate	239.40	53.9					
	$2NH_3\cdot H_2O$	Ammonia clathrate	191.90	58.4					
	$Kr\cdot 6H_2O$	Krypton clathrate	49.06	58.6					
	$NH_3\cdot H_2O$	Ammonia clathrate	35.05	80.3					

Figure 1. Calculated vacuum evaporation rates as a function of temperature for representative organic and inorganic compounds. The shaded areas show the range of evaporation rates calculated for the compounds shown in the figure legend. With the exception of sulfur, most inorganic volatiles are less volatile than water. Most simple organics and clathrates are also more volatile than water, whereas aromatic hydrocarbons, linear amides and carboxylic acids are less volatile than water. Very large molecules such as Fullerene (not shown on graph) are much less volatile than water.